

Download Experiment 8 Limiting Reactant Answers

In the reaction above, if we add more of the limiting reagent, reactant B, more of. However, the reactants for a reaction in an experiment are not necessarily a. Multiply the ratio by the limiting reactant's quantity in moles. The answer is the theoretical yield, in moles, of the desired product. In this example, the 25g of glucose equate to 0.139 moles of glucose. Write down the number of moles of your limiting reactant. The theoretical yield of an experiment is the amount of product created in perfect conditions. Limiting Reagents and Percentage Yield Worksheet: 1. Consider the reaction $\text{I}_2\text{O}_5(\text{g}) + 5\text{CO}(\text{g}) \rightarrow 5\text{CO}_2(\text{g}) + \text{I}_2(\text{g})$ a) 80.0 grams of iodine(V) oxide, I_2O_5 , reacts with 28.0 grams of carbon monoxide, CO .